

YEAR 12 A/ B – CHEMISTRY

WEEK 2 (29th March to 2nd April)

Work sent to the students through Whats app group / Google classroom / Zoom Learning Platform

Lesson Objective: Predict and justify the qualitative effect of a change in temperature, concentration or pressure on a homogeneous system in equilibrium.

Resources: Text book, Worksheet file, video, power point presentations.

<p>Tuesday – 6th, 7th & 8th period (Yr 12 A) Monday – 6th & 7th period (Yr 12 B)</p>	<p>Discuss the reasons for developing alternative fuels in terms of sustainability and reducing emissions, including the emission of CO₂ and its relationship to climate change.</p> <p>Know that many reactions are readily reversible and that they can reach a state of dynamic equilibrium in which:</p> <ol style="list-style-type: none">the rate of the forward reaction is equal to the rate of the backward reactionthe concentrations of reactants and products remain constant. <p>Write Answers to Textbook Ques in your notebook Complete worksheet file ques page</p>
<p>Wednesday – 7th & 8th period -Yr 12 A Tuesday – 1st & 2nd period (Yr 12 B)</p>	<p>Predict and justify the qualitative effect of a change in temperature, concentration or pressure on a homogeneous system in equilibrium.</p> <p>Evaluate data to explain the necessity, for many industrial processes, to reach a compromise between the yield and the rate of reaction.</p> <p>Write Answers to Text book Questions and Exam style questions in your notebook</p>

<p>Thursday – 7th period (Yr 12 A)</p> <p>Wednesday – 1st & 2nd period-Yr 12 B</p>	<p>Deduce an expression for K_c, for homogeneous and heterogeneous systems, in terms of equilibrium concentrations.</p> <p>Worksheet file questions page</p> <p>Discuss Exam style questions</p>
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