YEAR 13 A /B –CHEMISTRY

WEEK 3 (13th September to 17th September) 3 lessons for both batches(continuation from previous week)

Work sent to the students through Google classroom / Zoom Learning Platform

Resources: Text book, Worksheets, video, power point presentations.

Date	**Lesson**	**Topic**	**Mode of Teaching**	
13.09.2020 Sunday	4- **1 3B** **1- 13 A**	**Lesson Objective:** Calculation of value of equilibrium constant from given data. Calculation of units of Kc and Kp. **Success Criteria:** • writing equilibrium expression using the balanced chemical equations. • calculation of partial pressure from data.	**Zoom**	Teacher uses powerpoint presentation that contains interactive questions. Students solve the worksheet file questions and upload in the google classroom at end of the lesson
Sunday 13.09.2020 Wednesday 16.09.2020 Tuesday 15.09.2020 Wednesday 16.09.2020	2-13 A 4- 13A 2 - **13B** 2- **13 B**	**Lesson Objective:** Calculation of units of Kc and Kp. **Success Criteria** • Explain the steps to calculate value of Kp. • calculation of units for Kp using the given data.	**Zoom** **Zoom**	Teacher uses powerpoint presentation that contains interactive questions. Students solve the worksheet file questions and upload in the google classroom at end of the lesson

Homework : Solve the worksheet problems.

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Topic: Multidentate Ligands
Resources: Text book, Worksheet file, video, power point presentations.

Date	Lesson	Topic	Mode of Teaching	
13.09.2020 Sunday 16.09.2020 wednesday 13.09.2020 Sunday	4 **1 3A** 5 **13A** 5, 8 **13B**	**Learning objective:** Understand, in terms of the large positive increase in ΔSsystem, that the substitution of a monodentate ligand by a bidentate or multidentate ligand leads to a more stable complex ion. **Learning Outcome :** Explain how the substitution of small, uncharged ligands (such as H_2O) by larger, charged ligands (such as Cl−) leads to a change in coordination number and hence the shape of the complex. Explain how there is a large positive increase in ΔSsystem, when the substitution of a monodentate ligand by a bidentate or multidentate ligand takes place to form a more stable complex ion.	**Zoom**	Teacher uses power point to explain the different types of ligands and Students discuss and solve textbook questions. Students discuss and solve textbook questions.
15.09.2020 Tuesday 09.09.2020 Wednesday	1 **13B** 6 **13A**	**Learning objective**: Cite examples and identify bidentate ligands, such as ethylene diamine and multidentate ligands, such as $EDTA^{4-}$. **Learning Outcome :** Identify and explain the bidentate and multidentate ligands in complex formation.	**Zoom**	Students discuss and solve worksheet questions.

Home work: Solve worksheet file questions of transition metals.