YEAR 11 A/D/E – CHEMISTRY (Girls)

WEEK 7 (11th October to 15th October)

Work Sent to the students through Zoom Learning Platform / Google classroom

Topic:– SC12a: Dynamic equilibrium
SC15b: Factors affecting equilibrium

Resources: Text book, Worksheet, Board works power point

Date	Topic	
11.10.20 Sunday 8th period **Mode of Teaching:** Zoom	**Learning Objective:** Recall that chemical reactions are reversible, the use of the symbol ⇌ in equations and that the direction of some reversible reactions be altered by changing the reaction conditions. Explain what is meant by dynamic equilibrium. **Learning Outcome:** Define reversible reactions. Explain the significance ⇌ in equations. Cite **some** examples of reversible reactions. Define dynamic equilibrium. Suggest **some** examples of reactions in dynamic equilibrium.	Teacher uses power point presentation with interactive questions. Teacher uses textbook questions to assess the concept of reversible reactions.
12.10.20 Monday 4th period **Mode of Teaching:** Zoom	**Learning Objective:** Recall the conditions for the Haber process as: a) Temperature 450 °C b) pressure 200 atmospheres c) iron catalyst. Predict how the position of a dynamic equilibrium is affected by changes in: a) temperature b) pressure c) concentration. **Learning Outcome:** Analyse the conditions for the Haber process as: a) temperature 450 °C b) pressure 200 atmospheres c) iron catalyst Explain how the position of a dynamic equilibrium is affected by changes in: a) temperature b) pressure c) concentration	Teacher uses power point presentation with interactive questions. Use video of Haber process.
14.10.20 Wednesday 8th period	**Learning Objective:** Explain how, in industrial reactions, including the Haber process, conditions used are related to: a) the availability and cost of raw materials and energy supplies b) the control of temperature, pressure and catalyst used produce an	Teacher uses power point presentation with interactive questions and extra questions from past

Mode of Teaching: Zoom	acceptable yield in an acceptable time **Learning Outcome:** Analyze and evaluate the conditions for the Haber process as: a temperature 450 °C b pressure 200 atmospheres c iron catalyst	papers.
15.10.20 Thursday 5th and 6th Period	**Science GL Exam**	Class Teacher will conduct the GL exam and monitor the students on Zoom

HOMEWORK: Complete the textbook questions SC15b:Factors affecting equilibrium- page 122 - 123

YEAR 11 B/C/F – CHEMISTRY (Boys)

WEEK 7 (11th October to 15th October)
Work Sent to the students through Zoom Learning Platform / Google classroom
Topic:– SC12a: Dynamic Equilibrium
SC15b: Factors affecting equilibrium
Resources: Text book, Worksheet, Board works power point

Date	**Topic**	
11.10.20 Sunday 1st Period **Mode of Teaching:** Zoom	**Learning Objective:** Recall that chemical reactions are reversible, the use of the symbol ⇌ in equations and that the direction of some reversible reactions be altered by changing the reaction conditions. Explain what is meant by dynamic equilibrium. **Learning Outcome:** Define reversible reactions. Explain the significance ⇌ in equations. Cite **some** examples of reversible reactions. Define dynamic equilibrium. Suggest **some** examples of reactions in dynamic equilibrium.	Teacher uses power point presentation with interactive questions. Teacher uses textbook questions to assess the concept of reversible reactions.

11.10.20 Sunday 2nd Period **Mode of Teaching:** Zoom	**Learning Objective:** Recall the conditions for the Haber process as: a) temperature 450 °C b) pressure 200 atmospheres c) iron catalyst. **Learning Outcome:** Analyze the conditions for the Haber process as: a) temperature 450 °C b) pressure 200 atmospheres c) iron catalyst	Teacher uses power point presentation with interactive questions. Use video of Haber process.
12.10.20 Monday 3rd Period **Mode of Teaching:** Zoom	**Learning Objective:** Predict how the position of a dynamic equilibrium is affected by changes in: a) temperature b) pressure c) concentration. **Learning Outcome:** Explain how the position of a dynamic equilibrium is affected by changes in: a) temperature b) pressure c) concentration	Teacher uses power point presentation with interactive questions
13.10.20 Tuesday 7th Period **Mode of Teaching:** Zoom	**Learning Objective:** Explain how, in industrial reactions, including the Haber process, conditions used are related to: a) the availability and cost of raw materials and energy supplies b) the control of temperature, pressure and catalyst used produce an acceptable yield in an acceptable time **Learning Outcome:** Analyze and evaluate the conditions for the Haber process as: a temperature 450 °C b pressure 200 atmospheres c iron catalyst	Teacher uses power point presentation with interactive questions and extra questions from past papers.
15.10.20 Thursday 4th Period **Mode of**	**Learning Objective:** To answer the questions, on Factors affecting Equilibrium, in the worksheet. **Learning outcome:** Students will be able to reinforce the concepts learned in the previous lesson by answering the questions in the worksheet.	Worksheet assigned through GC. Instruction will be given in the GC to complete the worksheet.

Teaching: GC		

HOMEWORK: Complete the textbook questions SC15b:Factors affecting equilibrium- page 122 - 123

YEAR 11 G/H–CHEMISTRY (IGCSE)

WEEK 7 (11th Oct to 15th Oct)

Work Sent to the students through Google classroom/Zoom Learning Platform
Unit 3 – Topic: Reversible reactions and Equilibria
Resources: Text book, Worksheet, IGCSE science free lesson video, power point.

Date	Lesson	Topic	Mode of Teaching	
11.10.2020 Sunday	1 11**H** 6 **11G**	**Lesson Objective:** know that some reactions are reversible and this is indicated by the symbol ⇌ in equations **Learning Outcome**: Define reversible reactions. Discuss the significance ⇌ in equations. Cite **some** examples of reversible reactions.	**Zoom**	Teacher uses power point presentation with interactive questions and extra questions from past papers.

12.10.2020 Monday	2 **11H** 5 **11G**	**Lesson Objective:** know that a reversible reaction can reach dynamic equilibrium in a sealed container **Learning Outcome**: Define dynamic equilibrium. Suggest **some** examples of reactions in dynamic equilibrium. Discuss why dynamic equilibrium is possible only in a closed system/sealed container.	**Zoom**	Teacher uses a Ppt presentation/video that contains interactive questions.Teacher uses textbook questions to assess the concept of reversible reactions.
13.10.2020 Tuesday	3 **11H** 1 **11G**	**Lesson Objective:** know that the characteristics of a reaction at dynamic equilibrium are: • the forward and reverse reactions occur at the same rate • the concentrations of reactants and products remain constant **Learning Outcome:** State the characteristics of a reaction at dynamic equilibrium. Sketch a graph to show the characteristics of a reaction at dynamic equilibrium	**Zoom**	Teacher uses a PowerPoint presentation/ video to teach the concept of dynamic equilibrium.
	4**11H** 2 **11G**	**Lesson Objective:** Understand why a catalyst does not affect the position of equilibrium in a reversible reaction **Learning Outcome**: Discuss the effect of catalyst on the rate of a reaction at equilibrium. Give reasons why a catalyst does not affect the position of equilibrium in a reversible reaction.	**Zoom**	Instruction will be given in the class to complete the textbook and worksheet questions.
15.10. 2020 Thursday	5 **11H** 4 **11G**	**Lesson Objective:** Know the effect of changing either temperature or pressure on the position of equilibrium in a reversible reaction **Learning Outcome**: Explain how the position of a dynamic equilibrium is affected by changes in: a temperature b pressure c concentration **Predict** the change in conditions on the observation for reversible reaction	**GC**	Teacher sends PowerPoint presentation that contains interactive questions

HOMEWORK: Complete the textbook questions.