# YEAR 11 A/D/E – CHEMISTRY (Girls)

# WEEK 7 (11<sup>th</sup> October to 15<sup>th</sup> October)

### Work Sent to the students through Zoom Learning Platform / Google classroom

**Topic:** SC12a: Dynamic equilibrium

SC15b: Factors affecting equilibrium

**Resources:** Text book, Worksheet, Board works power point

Date	Topic			
11.10.20	Learning Objective:	Teacher uses		
Sunday 8 <sup>th</sup> period	Recall that chemical reactions are reversible, the use of the symbol  in equations and that the direction of some reversible reactions be  altered by changing the reaction conditions.	power point presentation with interactive questions.		
	Explain what is meant by dynamic equilibrium.	Teacher uses		
Mode of Teaching:	Learning Outcome:	textbook		
Zoom	Define reversible reactions.	questions to assess the		
	Explain the significance ≠ in equations.	concept of reversible		
	Cite <b>some</b> examples of reversible reactions.	reactions.		
	Define dynamic equilibrium.			
	Suggest <b>some</b> examples of reactions in dynamic equilibrium.			
	Learning Objective:	Teacher uses		
12.10.20	Recall the conditions for the Haber process as:	power point presentation		
Monday	a) Temperature 450 °C b) pressure 200 atmospheres c) iron catalyst.	with interactive		
4 <sup>th</sup> period	Predict how the position of a dynamic equilibrium is affected by changes in: a) temperature b) pressure c) concentration.	questions.		
Mode of	Learning Outcome:	Use video of		
Teaching:	Analyse the conditions for the Haber process as:	Haber process.		
Zoom	a) temperature 450 °C b) pressure 200 atmospheres c) iron catalyst			
	Explain how the position of a dynamic equilibrium is affected by changes in: a) temperature b) pressure c) concentration			
	Learning Objective:	Teacher uses		
14.10.20 Wednesday	Explain how, in industrial reactions, including the Haber process, conditions used are related to:	power point presentation with interactive		
8 <sup>th</sup> period	a) the availability and cost of raw materials and energy supplies	questions and extra questions		
_	b) the control of temperature, pressure and catalyst used produce an	from past		

Mode of	acceptable yield in an acceptable time	papers.
Teaching: Zoom	Learning Outcome:	
	Analyze and evaluate the conditions for the Haber process as:	
	a temperature 450 °C	
	b pressure 200 atmospheres	
	c iron catalyst	
15.10.20	Science GL Exam	Class Teacher will conduct the GL exam and
Thursday		monitor the
5 <sup>th</sup> and 6 <sup>th</sup> Period		students on Zoom

**HOMEWORK:** Complete the textbook questions SC15b:Factors affecting equilibrium- page 122 - 123

# YEAR 11 B/C/F - CHEMISTRY (Boys)

WEEK 7 (11<sup>th</sup> October to 15<sup>th</sup> October)

Work Sent to the students through Zoom Learning Platform / Google classroom

**Topic:** SC12a: Dynamic Equilibrium

SC15b: Factors affecting equilibrium

Resources: Text book, Worksheet, Board works power point

Date	Topic	
11.10.20	Learning Objective:	Teacher uses
Sunday 1 <sup>st</sup> Period	Recall that chemical reactions are reversible, the use of the symbol  in equations and that the direction of some reversible reactions be altered by changing the reaction conditions.	power point presentation with interactive questions.
	Explain what is meant by dynamic equilibrium.	
Mode of Teaching:	Learning Outcome:	Teacher uses
Zoom	Define reversible reactions.	textbook
	Explain the significance ≠ in equations.	questions to assess the concep
	Cite <b>some</b> examples of reversible reactions.	of reversible
	Define dynamic equilibrium.	reactions.
	Suggest <b>some</b> examples of reactions in dynamic equilibrium.	
		1

11.10.20	Learning Objective:	Teacher uses
Sunday	Recall the conditions for the Haber process as:	power point presentation with
2 <sup>nd</sup> Period	a) temperature 450 °C b) pressure 200 atmospheres c) iron catalyst.	interactive
	<b>Learning Outcome:</b>	questions.
Mode of Analyze the conditions for the Haber process as:		Use video of Haber process.
Teaching: Zoom	a) temperature 450 °C b) pressure 200 atmospheres c) iron catalyst	
12.10.20	Learning Objective:	Teacher uses
Monday	Predict how the position of a dynamic equilibrium is affected by	power point presentation with
3 <sup>rd</sup> Period	changes in: a) temperature b) pressure c) concentration.	interactive questions
		questions
	Learning Outcome:	
Mode of	Ü	
Teaching: Zoom	Explain how the position of a dynamic equilibrium is affected by changes in: a) temperature b) pressure c) concentration	
13.10.20	Learning Objective:	Teacher uses
Tuesday	Explain how, in industrial reactions, including the Haber process,	power point presentation with
7th D · 1	conditions used are related to:	interactive
7 <sup>th</sup> Period		11100100011
/ Period	a) the availability and cost of raw materials and energy supplies	questions and
Mode of	<ul><li>a) the availability and cost of raw materials and energy supplies</li><li>b) the control of temperature, pressure and catalyst used produce an acceptable yield in an acceptable time</li></ul>	
	b) the control of temperature, pressure and catalyst used produce an	questions and extra questions
Mode of Teaching:	b) the control of temperature, pressure and catalyst used produce an acceptable yield in an acceptable time	questions and extra questions
Mode of Teaching:	b) the control of temperature, pressure and catalyst used produce an acceptable yield in an acceptable time  Learning Outcome:	questions and extra questions
Mode of Teaching:	b) the control of temperature, pressure and catalyst used produce an acceptable yield in an acceptable time  Learning Outcome:  Analyze and evaluate the conditions for the Haber process as:	questions and extra questions
Mode of Teaching:	b) the control of temperature, pressure and catalyst used produce an acceptable yield in an acceptable time  Learning Outcome:  Analyze and evaluate the conditions for the Haber process as: a temperature 450 °C b pressure 200 atmospheres c iron catalyst  Learning Objective: To answer the questions, on Factors affecting	questions and extra questions from past papers.  Worksheet
Mode of Teaching: Zoom	b) the control of temperature, pressure and catalyst used produce an acceptable yield in an acceptable time  Learning Outcome:  Analyze and evaluate the conditions for the Haber process as:  a temperature 450 °C b pressure 200 atmospheres c iron catalyst	questions and extra questions from past papers.  Worksheet assigned
Mode of Teaching: Zoom	b) the control of temperature, pressure and catalyst used produce an acceptable yield in an acceptable time  Learning Outcome:  Analyze and evaluate the conditions for the Haber process as: a temperature 450 °C b pressure 200 atmospheres c iron catalyst  Learning Objective: To answer the questions, on Factors affecting Equilibrium, in the worksheet.	questions and extra questions from past papers.  Worksheet
Mode of Teaching: Zoom  15.10.20 Thursday	b) the control of temperature, pressure and catalyst used produce an acceptable yield in an acceptable time  Learning Outcome:  Analyze and evaluate the conditions for the Haber process as: a temperature 450 °C b pressure 200 atmospheres c iron catalyst  Learning Objective: To answer the questions, on Factors affecting Equilibrium, in the worksheet.  Learning outcome: Students will be able to reinforce the concepts	worksheet assigned through GC. Instruction will be given in the GC
Mode of Teaching: Zoom  15.10.20 Thursday	b) the control of temperature, pressure and catalyst used produce an acceptable yield in an acceptable time  Learning Outcome:  Analyze and evaluate the conditions for the Haber process as: a temperature 450 °C b pressure 200 atmospheres c iron catalyst  Learning Objective: To answer the questions, on Factors affecting Equilibrium, in the worksheet.	questions and extra questions from past papers.  Worksheet assigned through GC.  Instruction will be

Teaching:	
GC	

**HOMEWORK:** Complete the textbook questions SC15b:Factors affecting equilibrium- page 122 - 123

# YEAR 11 G/H-CHEMISTRY (IGCSE)

WEEK 7 (11<sup>th</sup> Oct to 15th Oct)

Work Sent to the students through Google classroom/Zoom Learning Platform

**Unit 3 – Topic:** Reversible reactions and Equilibria

Resources: Text book, Worksheet, IGCSE science free lesson video, power point.

Date	Lesson	Topic	Mode of Teaching	
11.10.2020 Sunday	1 11 <b>H</b> 6 <b>11G</b>	Lesson Objective: know that some reactions are reversible and this is indicated by the symbol ≠ in equations  Learning Outcome: Define reversible reactions.  Discuss the significance ≠ in equations. Cite some examples of reversible reactions.	Zoom	Teacher uses power point presentation with interactive questions and extra questions from past papers.

12.10.2020	2 <b>11H</b>	<b>Lesson Objective:</b> know that a reversible		Teacher uses a Ppt
12.10.2020	2 1111	reaction can reach dynamic equilibrium in		presentation/video
	5 11G	a sealed container	Zoom	1 *
				that contains
Monday		Learning Outcome:		interactive
		Define dynamic equilibrium.		questions.Teacher
		Suggest <b>some</b> examples of reactions in		uses textbook
		dynamic equilibrium.		questions to
		Discuss why dynamic equilibrium is		assess the concept
		possible only in a closed system/sealed		of reversible
		container.		reactions.
				reactions.
13.10.2020	3 <b>11H</b>	Lesson Objective:		Teacher uses a
		know that the characteristics of a reaction		PowerPoint
	1 <b>11G</b>	at dynamic equilibrium are:	Zoom	presentation/
		• the forward and reverse reactions occur		-
Tuesday		at the same rate		video to teach the
		• the concentrations of reactants and		concept of
		products remain constant		dynamic
				equilibrium.
		Learning Outcome:		
		State the characteristics of a reaction at		
		dynamic equilibrium.		
		Sketch a graph to show the characteristics		
		of a reaction at dynamic equilibrium		
	44477	T 01: 4:	77	T ( '' '111
	411H	Lesson Objective:	Zoom	Instruction will be
	2 11G	Understand why a catalyst does not affect the position of equilibrium in a reversible		given in the class
	2 110	reaction		to complete the
		Learning Outcome:		textbook and
		Discuss the effect of catalyst on the rate of		worksheet
		a reaction at equilibrium.		questions.
		Give reasons why a catalyst does not		-
		affect the position of equilibrium in a		
		reversible reaction.		
15.10. 2020	5 <b>11H</b>	Lesson Objective:	GC	Teacher sends
		Know the effect of changing either		PowerPoint
	4 <b>11G</b>	temperature or pressure on the position of		presentation that
TOTAL 1		equilibrium in a reversible reaction		contains
Thursday		<b>Learning Outcome</b> :		interactive
		Explain how the position of a dynamic		
		equilibrium is affected by changes in:		questions
		a temperature		
		b pressure		
		c concentration		
		<b>Predict</b> the change in conditions on the		
		observation for reversible reaction		

**HOMEWORK:** Complete the textbook questions.