

## YEAR 10 B/C/F–CHEMISTRY (Boys)

WEEK 9 (25<sup>th</sup> Oct to 28<sup>th</sup> Oct)

Work Sent to the students through Google classroom

Topic: Bases and salts

Resources: Text book, Worksheet, power point.

Date	Lesson	Topic	Mode of Teaching	
25/10/2020 Sunday	0	<p><b>Learning Objective:</b></p> <ol style="list-style-type: none"><li>1. Recall that a base is any substance that reacts with an acid to form a salt and water only.</li><li>2. Explain the general reactions of aqueous solutions of acids with metal oxides to produce salts.</li><li>3. Write balanced equations, including the use of the state symbols (s), (l), (g) and (aq).</li></ol> <p><b>Learning Outcome:</b></p> <ul style="list-style-type: none"><li>• Define base.</li><li>• Describe what happens when an acid reacts with a metal oxide.</li><li>• Write symbol equations for the reactions of acids and metal oxides.</li><li>• Explain what happens during a neutralization reaction.</li></ul>	<b>Zoom</b>	Teacher uses power point presentation to explain the reactions of aqueous solutions of acids with metal oxides to produce salts.
26/10/2020 Monday	1&2	<p><b>Learning Objective:</b></p> <ol style="list-style-type: none"><li>1. Explain why, if soluble salts are prepared from an acid and an insoluble reactant:<ol style="list-style-type: none"><li>a excess of the reactant is added</li><li>b the excess reactant is removed</li><li>c the solution remaining is only salt and water.</li></ol></li><li>2. Describe the steps involved in preparing a soluble salt from an acid and an insoluble reactant.</li></ol> <p><b>Learning Outcome:</b></p> <ul style="list-style-type: none"><li>• Explain why an excess of insoluble reactant is used when preparing a soluble salt.</li><li>• Explain why the excess reactant is removed when preparing a soluble salt.</li><li>• Explain why the remaining solution contains only</li></ul>	<b>Zoom</b>	Teacher uses power point presentation that contains the steps involved in preparing a soluble salt from an acid and an insoluble reactant..

		a salt and water, when preparing a soluble salt from an acid and an insoluble reactant.		
28/10/2020 Wednesday	4	<p><b>Learning Objective:</b></p> <p>To answer the questions, on bases and salts, in the worksheet.</p> <p><b>Learning outcome:</b></p> <p>Students will be able to reinforce the concepts learned in the previous lesson by answering the questions in the worksheet.</p>	GC	Instruction will be given in the Google classroom to complete the Worksheet.

Home work: Solve S1 and E1 questions :SC8c(Pg57)

### **YEAR 10 A/D/E–CHEMISTRY (girls)**

**WEEK 9 (25<sup>th</sup> Oct to 28<sup>th</sup> Oct)**

**Work Sent to the students through Google classroom**

**Topic:** Bases and salts.

**Resources:** Text book, Worksheet, power point.

Date	Lesson	Topic	Mode of Teaching	
25/10/2020 Sunday	3	<p><b>Learning Objective:</b></p> <p><b>Learning Objective:</b></p> <ol style="list-style-type: none"> <li>1. Recall that a base is any substance that reacts with an acid to form a salt and water only.</li> <li>2. Explain the general reactions of aqueous solutions of acids with metal oxides to produce salts.</li> <li>3. Write balanced equations, including the use of the state symbols (s), (l), (g) and (aq).</li> </ol> <p><b>Learning Outcome:</b></p> <ul style="list-style-type: none"> <li>•Define base.</li> <li>•Describe what happens when an acid reacts with a metal oxide.</li> <li>•Write symbol equations for the reactions of acids and metal oxides.</li> </ul>	Zoom	Teacher uses power point presentation to explain the reactions of aqueous solutions of acids with metal oxides to produce salts.

		Explain what happens during a neutralization reaction		
28/10/2020 Wednesday	3	<p><b>Learning Objective:</b></p> <p>1.Explain why, if soluble salts are prepared from an acid and an insoluble reactant:</p> <p>a excess of the reactant is added</p> <p>b the excess reactant is removed</p> <p>c the solution remaining is only salt and water.</p> <p>2.Describe the steps involved in preparing a soluble salt from an acid and an insoluble reactant.</p> <p><b>Learning Outcome:</b></p> <ul style="list-style-type: none"> <li>• Explain why an excess of insoluble reactant is used when preparing a soluble salt.</li> <li>• Explain why the excess reactant is removed when preparing a soluble salt.</li> <li>• Explain why the remaining solution contains only a salt and water, when preparing a soluble salt from an acid and an insoluble reactant.</li> </ul>	<b>Zoom</b>	Teacher uses power point presentation that contains the steps involved in preparing a soluble salt from an acid and an insoluble reactant..

Home work: Solve S1 and E1 questions :SC8c(Pg57)