

## YEAR 10 A/D/E-CHEMISTRY (girls)

**WEEK 6 (4<sup>th</sup> Oct to 8<sup>th</sup> Oct)**

**Work Sent to the students through Google classroom**

**Topic:** Moles

**Resources:** Text book, Worksheet, power point.

<b>Date</b>	<b>Lesson</b>	<b>Topic</b>	<b>Mode of Teaching</b>	
4/10/2020 Sunday	3	<p><b>Learning Objective:</b> To reinforce the role of limiting reagent in deciding the amount of product formed in a reaction.</p> <p><b>Learning outcome:</b></p> <ul style="list-style-type: none"> <li>• What controls the mass of product formed in a reaction?</li> <li>• Reason out why mass of product formed is controlled by the limiting reagent.</li> <li>• Differentiate between excess and limiting reagent.</li> </ul>	<b>Zoom</b>	Teacher uses powerpoint presentation to reinforce the role of limiting reagent in deciding the amount of product formed in a reaction.
7/10/2020 Wednesday	3	<p><b>Learning Objective:</b> Deduce the balanced equation for a reaction from the masses of reactants and/or products.</p> <p><b>Learning Outcome:</b></p> <ul style="list-style-type: none"> <li>• Calculate the moles of reactants and products.</li> <li>• Work out the equation for a chemical reaction using the masses of the reactants and products.</li> <li>• Write a balanced chemical equation for a reaction.</li> </ul>	<b>Zoom</b>	Teacher uses powerpoint presentation that contains the steps to deduce the stoichiometry of a reaction.
8/10/2020 Thursday	2  3	<p><b>Learning Objective:</b> To reinforce the calculation of masses of reactants and products from balanced equations, given the mass of one substance.</p> <p><b>Learning Outcome:</b></p> <ul style="list-style-type: none"> <li>• Calculate the mass of product formed from a given mass of reactant, using a balanced equation</li> <li>• Calculate the mass of a reactant needed to produce a given amount of product, using a balanced equation.</li> </ul>	<b>Zoom</b>  <b>GC</b>	Teacher uses powerpoint presentation that contains interactive questions to reinforce the calculation of masses of reactants and products Instruction will be given in the Google classroom to complete the Worksheet.

Home work: Solve textbook questions 5&6:SC9c(Pg77)

## YEAR 10 B/C/F–CHEMISTRY (Boys)

WEEK 6 (4<sup>th</sup> Oct to 8<sup>th</sup> Oct)

Work Sent to the students through Google classroom

Topic: Moles

Resources: Text book, Worksheet, power point.

Date	Lesson	Topic	Mode of Teaching	
4/10/2020 Sunday	0	<b>Learning Objective:</b> To reinforce the role of limiting reagent in deciding the amount of product formed in a reaction. <b>Learning outcome:</b> <ul style="list-style-type: none"><li>• What controls the mass of product formed in a reaction?</li><li>• Reason out why mass of product formed is controlled by the limiting reagent.</li><li>• Differentiate between excess and limiting reagent.</li></ul>	<b>Zoom</b>	Teacher uses powerpoint presentation to reinforce the role of limiting reagent in deciding the amount of product formed in a reaction.
5/10/2020 Monday	1&2	<b>Learning Objective:</b> Deduce the balanced equation for a reaction from the masses of reactants and/or products. <b>Learning Outcome:</b> <ul style="list-style-type: none"><li>• Calculate the moles of reactants and products.</li><li>• Work out the equation for a chemical reaction using the masses of the reactants and products.</li></ul> Write a balanced chemical equation for a reaction.	<b>Zoom</b>	Teacher uses powerpoint presentation that contains the steps to deduce the stoichiometry of a reaction.
7/10/2020 Wednesday	4	<b>Learning Objective:</b> To reinforce the calculation of masses of reactants and products from balanced equations, given the mass of one substance. <b>Learning Outcome:</b> <ul style="list-style-type: none"><li>• Calculate the mass of product formed from a given mass of reactant, using a balanced equation</li><li>• Calculate the mass of a reactant needed to produce a given amount of product, using a balanced equation.</li></ul>	<b>GC</b>	Instruction will be given in the Google classroom to complete the Worksheet.

Home work:Solve textbook questions 5&6:SC9c(Pg77)

